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Thermodynamics and Energy. Thermodynamics can be defined as the study of energy, energy transformations and its relation to matter. The analysis of thermal systems is achieved through the application of the governing conservation equations, namely Conservation of Mass, Conservation of Energy (1st law of thermodynamics), the 2nd law of thermodynamics and the property relations.

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The difference between thermal energy and heat is that thermal energy is not being transferred, but remains part of the system's internal energy (kinetic energy of its internal particles); Instead, heat is energy in the transfer. Heat energy is transferred from a hot-system to another cold system according to the 2nd law of thermodynamics. An example of the difference is a burning candle. A candle generates thermal energy. While it is burning, it dissipates its thermal energy into the ...

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Thermodynamics Teacher and Student Guides. This is a hands-on laboratory unit from the U.S. Department of Energy exploring the concepts of heat and movement. Teachers set up six laboratory stations that will introduce students to the basic concepts of thermodynamics, including atomic structure, atomic and molecular motion, states of matter, heat transfer, thermal expansion, specific heat, and heats of fusion and vaporization.

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Thermodynamics is from the Greek root meaning movement of heat. "The study of heat and its transformation into different forms of energy." Coined by Lord Kelvin is the study of thermodynamics concerned primarily with microscopic processes or with macroscopic ones? Why? Macroscopic, because atoms were unknown at the time thermodynamics was created.

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The second law of thermodynamics. Heat does not flow spontaneously from a colder region to a hotter region, or, equivalently, heat at a given temperature cannot be converted entirely into work. Consequently, the entropy of a closed system, or heat energy per unit temperature, increases over time toward some maximum value. Thus, all closed systems tend toward an equilibrium state in which entropy is at a maximum and no energy is available to do useful work.

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Definition of temperature. The definition of thermodynamic temperature T is a function of the change in the system's entropy S under reversible heat transfer Q rev: = . Entropy being a state function, the integral of dS over any cyclical process is zero. For a system in which the entropy is purely a function of the system's energy E, the temperature can be defined as:

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