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and work effects and relations between heat and work. Some early thermodynamics problems were for very practical problems. For example, in a steam engine heat

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**is supplied to water to
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generic heat capacity c
105 bT Cp = a +
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The heat capacities of
various elements and**

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compounds are $C_p = a + bT + cT^2$
 $a = 21.30$, $b = 8.54 \times 10^{-5}$,
 $c = 1.51 \times 10^{-8}$...

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thermodynamics has much wider applicability. In material science, one is normally not that interested in heat and work, but interested more the state of matter

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and how things might change when mixed, heated, pressurized, etc. Some important effects are chemical reactions (such as oxidation), formation of

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**solutions, phase
transformations.**

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**Solution the first law
as $w = q - \Delta U$; thus
 $q = \Delta H$; $w = \Delta H - \Delta U$; 4.**

Isothermal Process

**Because U is a function
only of T for an ideal
gas, $\Delta U = \Delta H = 0$ for an**

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**isothermal process.
These results also
follow from the general
results by using $DT =$
 $D(PV) = 0$ for an
isothermal process.**

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**Define a trivial
function $\rho = a + z$ in
equation window, (2)
select new parametric**

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**table from Tables, and
type the data in a two-
column table, (3) select
Plot and plot the data,
and (4) select plot and
click on “curve fit” to
get. curve fit window.**

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relations between heat and work. Some early thermodynamics problems were for very practical problems. For example, in a steam engine heat is supplied to water to

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create steam.

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generic heat capacity c

$$105 \text{ bT } C_p = a +$$

$$bT +$$

$$cT^2 ; T_2 103$$

The heat capacities of

various elements and

compounds are $C_{p, \text{Ag}} = C_p$

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**ê. 8a Ø 21.30, b Ø 8.54,
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wider applicability. In material science, one is normally not that interested in heat and work, but interested more the state of matter and how things might

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**change when mixed,
heated, pressurized,
etc. Some important
effects are chemical
reactions (such as
oxidation), formation of
solutions, phase**

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as $w = q - \Delta U$; thus
 $q = \Delta H$; $w = -\Delta H$; 4.

Isothermal Process

Because U is a function only of T for an ideal gas, $\Delta U = \Delta H = 0$ for an isothermal process.

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These results also follow from the general results by using $DT = D(PV) = 0$ for an isothermal process.

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function $\rho = a + z$ in

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table from Tables, and

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type the data in a two-column table, (3) select Plot and plot the data, and (4) select plot and click on “curve fit” to get. curve fit window.

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