

Chemistry Chapter 12 Stoichiometry Worksheet

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problem, the first step is always to calculate the number of moles of each reactant present. In this case, we are giv
Cr 2 O 7 in 1 mL of solution, which we can use to calculate the number of moles of K 2 Cr ...

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equation?. $Fe_2O_3(s) + 3SO_3(g) \rightarrow Fe_2(SO_4)_3$. Solution. We use the balanced chemical equation to construct
factor between Fe 2 O 3 and SO 3. The number of molecules of Fe 2 O 3 goes on the bottom of our conversion fact
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Practice: 1. In the reaction: $4Li(s) + O_2(g) \rightarrow 2Li_2O(s)$ a. what is the product? $Li_2O(s)$ b. what are the reactant
c. what does the "(s)" after the formula of lithium oxide signify? phase ...

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