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*Chem 12 Notes On  
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Chemistry

~~Chemistry Notes | how to  
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*Naming Acids Introduction*

*Ka Kb Kw pH pOH pKa pKb*

*H<sup>+</sup> OH<sup>-</sup> Calculations - Acids*

*\u0026 Bases, Buffer*

*Solutions , Chemistry*

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*Review ACIDS BASES*

*\u0026amp; SALTS-FULL*

*CHAPTER || CLASS 10*

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*Chemistry: Acids Bases and*

*Salts pH, pOH, H<sub>3</sub>O<sup>+</sup>, OH<sup>-</sup>,*

*K<sub>w</sub>, K<sub>a</sub>, K<sub>b</sub>, pK<sub>a</sub>, and pK<sub>b</sub>*

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*Basic Calculations -Acids  
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*Problems IGCSE*

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~~[Syllabus 8] - Acids And~~

~~Bases Acid-Base Reactions  
in Solution: Crash Course~~

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*Chemistry #8 Class 10th  
chapter 2 Acid base and salt  
science PH numericals  
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Bases and Salts | Class 7  
Science Sprint for Final*

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*Exams | Chapter 5*

*@Vedantu Young Wonders*

*HCL - Hydrochloric Acid ||*

*ICSE CLASS 10*

*CHEMISTRY || Acids Bases*

*and Salts Class 10 Science*

*Chemistry CBSE NCERT*

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Periodic Table: Crash  
Course Chemistry #4 ~~Acids  
Bases and Salts Le  
Chatelier's Principle of~~*

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~~Chemical Equilibrium – Basic  
Introduction Calculating pH,  
pOH, [H<sup>+</sup>], [H<sub>3</sub>O<sup>+</sup>], [OH<sup>-</sup>] of  
Acids and Bases – Practice  
How do Physics Wallah Earn  
? Acids and Bases, pH and  
pOH How to Speak~~



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*ACID BASES AND SALTS |  
FULL CHAPTER | CLASS 10  
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01 : ACIDS : CBSE / ICSE  
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Chemistry

*Chem 12 Notes On Acids  
Chemistry 12 Notes on Unit  
4 - Acids, Bases and Salts  
Chemistry 12 -Unit 4 -Notes  
Page 2 - Any acid (weak or  
strong) could have high or  
low concentration. Weak &*

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### Chemistry

*Strong refers to % ionization. Concentration is the moles of acid dissolved per litre. Eg.) 10.0 M HCl is conc. and strong  $[H_3O^+] = 10.0 M$  0.001 M HCl is dilute and strong  $[H$*

# Where To Download Chem 12 Notes On Acids Bases Sss Chemistry

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*Chem 12- Notes on Acids &  
Bases - SSS Chemistry  
Acids from HClO<sub>4</sub> to H<sub>2</sub>SO<sub>4</sub>  
are 100% ionized in water .  
only solvent used in Chem*

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Chemistry

*12 (and most Chemistry) so even though  $\text{HClO}_4$  is above  $\text{HCl}$  on the chart, it is no more acidic in a water solution.  $\text{H}_3\text{O}^+$  is the strongest acid that can exist in an undissociated form in*



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Chemistry

*water solution. all stronger  
acids ionize to form  $H_3O^+$*

---

*Chem 12- Notes on Acids &  
Bases - SSS Chemistry*

*In this chapter, we will focus*

## Where To Download Chem 12 Notes On Acids Bases Sss

### Chemistry

*on acids and bases and their  
chemistry. 12.1:*

*Introduction Certain  
household chemicals, such  
as some brands of cleanser,  
can be very concentrated  
bases, which makes them*

Where To Download Chem 12  
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Chemistry

*among the most potentially hazardous substances found around the home; if spilled on the skin, the strong caustic compound can immediately remove  $H^+$  ions from the flesh, resulting in*

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Chemistry  
*chemical burns.*

---

*12: Acids and Bases -  
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Bases Chemistry 12 Notes*  
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Chemistry

*on Unit 4 - Acids, Bases and  
Salts Chemistry 12 -Unit 4  
-Notes Page 2 - Any acid  
(weak or strong) could have  
high or low concentration.  
Weak & Strong à refers to %  
ionization. Concentration à*

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Chemistry

*the moles of acid dissolved  
per litre. Eg.) 10.0 M HCl à  
conc. and strong  $[H_3O^+] =$*

---

*Chem 12 Notes On Acids  
Bases Sss Chemistry*

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Chemistry

*Chemistry 12 Notes on Unit  
4 - Acids, Bases and Salts*

*Chemistry 12 -Unit 4 -Notes*

*Page 2 - Any acid (weak or  
strong) could have high or  
low concentration. Weak &  
Strong à refers to %*

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Chemistry

*ionization. Concentration à  
the moles of acid dissolved  
per litre. Eg.) 10.0 M HCl à  
conc. and strong  $[H_3O^+] =$   
10.0 M 0.001 M HCl à dilute  
and strong  $[H$  Chem 12-  
Notes on Acids & Bases -*



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Chemistry

*SSS Chemistry 12.4: Acid-Base Titrations A titration is the quantitative reaction of an acid and a base.*

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*Chem 12 Notes On Acids*

*Page 25/114*

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Chemistry

*Bases Sss Chemistry D  
Colgur*

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Salts Chemistry 12 -Unit 4*

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*-Notes Page 2 - Any acid (weak or strong) could have high or low concentration. Weak & Strong à refers to % ionization. Concentration à the moles of acid dissolved per litre. Eg.) 10.0 M HCl à*

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*conc.*

---

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Chemistry

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*DETAILED NOTES ON  
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AND ACID BASE*

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*Strong Acids and Bases*

*VIDEO. Weak Acids VIDEO.*

*Weak Bases VIDEO*

*Page 30/114*

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*Chemistry 12*

*Acids and Bases - Section 14  
of General Chemistry Notes  
is 23 pages in length (page  
14-1 through page 14-23)*

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## Chemistry

*and covers ALL you'll need  
to know on the following  
lecture/textbook topics:.*  
*SECTION 14 - Acids and  
Bases 14-1 -- Acid-Base  
Chemistry · Arrhenius Acids  
and Bases · Bronsted-Lowry*



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*Acids and Bases · Proton  
(H<sup>+</sup>) Donors and Proton  
(H<sup>+</sup>) Acceptors*

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*Chemistry Notes | Chemistry  
Pdf -- Acids, Bases, and the*

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...

*1. The properties of acids and bases 2. pH and pOH calculations 3. Definitions of acids and bases 4.*

*Neutralization reactions.*

*NOTES: Properties of acids :*

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- *Form hydrogen ions ( $H^+$  or protons) in solution*
- *Form hydronium (a water molecule bonded to a hydrogen ion shown as  $H_3O^+$  or  $H^+$ ).*

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CHAPTERS 20 AND 21  
Acids and Bases ...*

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*open all notes for that unit.*

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*Kinetics Unit 2 Equilibrium*

*Unit 3 Solubility Equilibrium*

*Unit 4 Acids, Bases, and*

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*Salts Unit 5*

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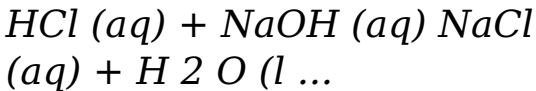
*Chemistry 12 Unit IV -  
Acids, Bases and Salts Notes  
IV.1 - The Arrhenius Theory  
Of Acids and Bases • Acid:  
any substance which  
releases  $H^+ (aq)$  in water. •*

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### Chemistry

*Base: any substance which releases  $\text{OH}^-(\text{aq})$  in water. •*

*Salt: is the neutralization product which results when an acid and a base react:*





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## *Chemistry 12*

*Acids are classified as  
Organic Acids and Mineral  
Acids. Acids which are  
derived from plants and*

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### Chemistry

*animals, they are known as Organic Acids. For Example, Citric Acid from fruit.*

*Mineral acids are inorganic acids such as Sulphuric Acid. They are dangerous to be used, so need more*

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*precautions. Acids are also classified as Strong Acids or Weak Acids. Strong acid is an acid, that completely dissociates into ions in aqueous solutions.*

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Chapter 2 - Acids, bases and*

*...*

*Acid-Base Titrations.  
Indicators. Titration Curves  
& Solving AB Titration*

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Chemistry

*Problems (Hebden  
summary) Acid Rain. Notes  
from demo: Lab Instructions  
Acid-Base Titration  
Suggested Questions to prep  
for written component of  
Acid/Base Unit Exam:*

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*Hebden Q #: 67, 69.e), 99,  
101, 104, 113, 115, 118,  
120, 125.e), 126, 130, 139,  
140*

---

*Unit 4 Acids, Bases, and*

*Page 46/114*

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Chemistry

*Salts | Grade 11 & 12 Notes  
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Curves, pKa and pH)*

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*CONCISE A\* A Level  
Chemistry Edexcel Chapter  
12 Notes ...  
Phosphoric acid (H<sub>3</sub>PO<sub>4</sub>)  
etc. Chemical Properties of*  
Page 48/114



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*Acid: (i) Reaction of acids with metal: Acids give hydrogen gas along with respective salt when they react with a metal. Metal + Acid  $\rightarrow$  Salt + Hydrogen*

*Examples: Hydrogen gas*

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*and zinc chloride are formed  
when hydrochloric acid  
reacts with zinc metal.*

---

*Acids Bases and Salts Class  
10 Notes Science Chapter 2*

*Page 50/114*

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*I) Weak Acid Equilibrium*

*and  $K_a$ ; II) Weak Base*

*Equilibrium and  $K_b$ ; III)*

*Writing Formula*

*(Molecular), Complete Ionic*

*and Net Ionic Equations for*

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Chemistry  
*Acid/Base ...*

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Notes Chapter - 2 Acids,  
Bases and Salts ACIDS: •*

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*These are the substances which have sour in taste. • They turns blue litmus solution to red. • They give H ions in aqueous solution*  
*Examples of Acids : - HCl - Hydrochloric Acid BASES •*

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### Chemistry

*These substances are bitter in taste. • They turns red litmus solution to blue.*

---

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*K<sub>a</sub> K<sub>b</sub> K<sub>w</sub> pH pOH pK<sub>a</sub> pK<sub>b</sub>  
H<sup>+</sup> OH<sup>-</sup> Calculations - Acids*

*\u0026 Bases, Buffer*

*Solutions , Chemistry*

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Chemistry: Acids Bases and  
Salts pH, pOH, H<sub>3</sub>O<sup>+</sup>, OH<sup>-</sup>,  
K<sub>w</sub>, K<sub>a</sub>, K<sub>b</sub>, pK<sub>a</sub>, and pK<sub>b</sub>  
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Problems ~~IGCSE~~*

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Introduction Calculating pH,  
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*Chemistry 12 Notes on Unit*

*4 - Acids, Bases and Salts*

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*Chemistry 12 -Unit 4 -Notes*

*Page 2 - Any acid (weak or strong) could have high or low concentration. Weak & Strong à refers to % ionization. Concentration à the moles of acid dissolved*

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Chemistry

*per litre. Eg.) 10.0 M HCl à conc. and strong [H<sub>3</sub>O<sup>+</sup>] = 10.0 M 0.001 M HCl à dilute and strong [H*

---

*Chem 12- Notes on Acids &*

*Page 71/114*

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Chemistry

*Bases - SSS Chemistry*

*Acids from HClO<sub>4</sub> to H<sub>2</sub>SO<sub>4</sub>*

*are 100% ionized in water .*

*only solvent used in Chem*

*12 (and most Chemistry) so*

*even though HClO<sub>4</sub> is above*

*HCl on the chart, it is no*



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Chemistry

*more acidic in a water solution.  $H_3O^+$  is the strongest acid that can exist in an undissociated form in water solution. all stronger acids ionize to form  $H_3O^+$*

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## *Chem 12- Notes on Acids & Bases - SSS Chemistry*

*In this chapter, we will focus  
on acids and bases and their  
chemistry. 12.1:*

*Introduction Certain*

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### Chemistry

*household chemicals, such as some brands of cleanser, can be very concentrated bases, which makes them among the most potentially hazardous substances found around the home; if spilled*

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*on the skin, the strong caustic compound can immediately remove  $H^+$  ions from the flesh, resulting in chemical burns.*

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*on Unit 4 - Acids, Bases and  
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*-Notes Page 2 - Any acid*

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*(weak or strong) could have high or low concentration.*

*Weak & Strong à refers to % ionization. Concentration à the moles of acid dissolved per litre. Eg.) 10.0 M HCl à conc. and strong  $[H_3O^+] =$*

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*conc. and strong [H<sub>3</sub>O<sup>+</sup>] =  
10.0 M 0.001 M HCl à dilute  
and strong [H<sup>+</sup> Chem 12-  
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SSS Chemistry 12.4: Acid-  
Base Titrations A titration is  
the quantitative reaction of*

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Chemistry  
*an acid and a base.*

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Salts Chemistry 12 -Unit 4  
-Notes Page 2 - Any acid  
(weak or strong) could have  
high or low concentration.*

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*Weak & Strong à refers to % ionization. Concentration à the moles of acid dissolved per litre. Eg.) 10.0 M HCl à conc.*

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*AND ACID BASE*

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*...*

*1. The properties of acids  
and bases 2. pH and pOH*

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Chemistry

*calculations 3. Definitions of acids and bases 4.*

*Neutralization reactions.*

*NOTES: Properties of acids :*

- Form hydrogen ions ( $H^+$  or protons) in solution*
- Form hydronium (a water*

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Chemistry

*molecule bonded to a  
hydrogen ion shown as H.  
3O+ or H+. H.*

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CHAPTERS 20 AND 21***

*Page 92/114*

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Chemistry

*Acids and Bases ...*

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*Prescribed Learning*

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Chemistry

*Notes Unit 1 Reaction*

*Kinetics Unit 2 Equilibrium*

*Unit 3 Solubility Equilibrium*

*Unit 4 Acids, Bases, and  
Salts Unit 5*

*Electrochemistry Link to  
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*Exams...*

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*Chemistry 12 | Grade 11 &  
12 Notes  
Chemistry 12 Unit IV -  
Acids, Bases and Salts Notes*  
*Page 95/114*

## Where To Download Chem 12 Notes On Acids Bases Sss

### Chemistry

*IV.1 - The Arrhenius Theory  
Of Acids and Bases*

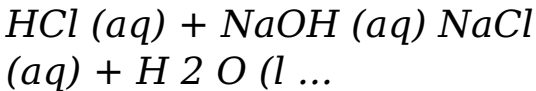
- *Acid: any substance which releases  $H^+(aq)$  in water.*
- *Base: any substance which releases  $OH^-(aq)$  in water.*
- *Salt: is the neutralization*



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Chemistry

*product which results when  
an acid and a base react:*



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*Chemistry 12*

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## Where To Download Chem 12 Notes On Acids Bases Sss Chemistry

*Acids are classified as Organic Acids and Mineral Acids. Acids which are derived from plants and animals, they are known as Organic Acids. For Example, Citric Acid from fruit.*

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*Mineral acids are inorganic acids such as Sulphuric Acid. They are dangerous to be used, so need more precautions. Acids are also classified as Strong Acids or Weak Acids. Strong acid is*

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*an acid, that completely dissociates into ions in aqueous solutions.*

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*Revision Notes for Science  
Chapter 2 - Acids, bases and*  
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Chemistry

...

*Acid-Base Titrations.  
Indicators. Titration Curves  
& Solving AB Titration  
Problems (Hebden  
summary) Acid Rain. Notes  
from demo: Lab Instructions*

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Chemistry

*Acid-Base Titration*

*Suggested Questions to prep  
for written component of  
Acid/Base Unit Exam:*

*Hebden Q #: 67, 69.e), 99,  
101, 104, 113, 115, 118,  
120, 125.e), 126, 130, 139,*

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140

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*Unit 4 Acids, Bases, and  
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Comprehensive notes on the  
Edexcel A Level Chemistry*

*Page 103/114*

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Chemistry

*Specification - Chapter 12.  
Notes on Acids and Bases,  
Buffer Solutions, Titration  
Curves,  $pK_a$  and  $pH$ )*

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**CONCISE A\* A Level**

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Chemistry

*Chemistry Edexcel Chapter  
12 Notes ...*

*Phosphoric acid ( $H_3PO_4$ )  
etc. Chemical Properties of  
Acid: (i) Reaction of acids  
with metal: Acids give  
hydrogen gas along with*

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Chemistry

*respective salt when they react with a metal. Metal + Acid → Salt + Hydrogen*  
*Examples: Hydrogen gas and zinc chloride are formed when hydrochloric acid reacts with zinc metal.*

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*Acids Bases and Salts Class  
10 Notes Science Chapter 2*

...

*CHemistry 12. Home Math 9  
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Chemistry

10 Chemistry 11 Chemistry

12 Class Grades Important

Documents: chem\_12\_course

\_outline\_zukowski.pdf: File

Size: 377 kb ... Class Notes:

I) Weak Acid Equilibrium

and  $K_a$ ; II) Weak Base

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Chemistry

*Equilibrium and  $K_b$ ; III)*

*Writing Formula*

*(Molecular), Complete Ionic  
and Net Ionic Equations for  
Acid/Base ...*

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Chemistry

*Chemistry 12 - Miss*

*Zukowski's Class*

*Notes Chapter - 2 Acids,*

*Bases and Salts ACIDS: •*

*These are the substances  
which have sour in taste. •*

*They turns blue litmus*

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*solution to red. • They give  
H ions in aqueous solution*

*Examples of Acids : - HCl -*

*Hydrochloric Acid BASES •*

*These substances are bitter  
in taste. • They turns red  
litmus solution to blue.*

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*Chapter - 2, Acids Bases and  
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