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## ***Chapter 9 Mixed Review Stoichiometry Answers***

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CHAPTER 9 REVIEW  
Stoichiometry SECTION 3  
PROBLEMS Write the answer  
on the line to the left. Show all  
your work in the space**

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- provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.
2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: 
$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$$

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#### Stoichiometry SECTION 3

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

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2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$  ...

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#### CHAPTER 9 REVIEW

#### Stoichiometry CHAPTER 9

#### REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER

Answer the following questions in the space provided. 1. Given the

following equation:  $C_3H_4(g) + x O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient  $x$  in this equation? b. What is the molar

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mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ?

## ~~Chapter 9 Review Stoichiometry Answers Section 2~~

### CHAPTER 9 REVIEW

#### Stoichiometry MIXED REVIEW

#### SHORT ANSWER Answer the

following questions in the

space provided. 1. Given the

following equation:  $C_3H_4(g)$

$+ xO_2(g) \rightarrow 3CO_2(g) + 2H$

$2O(g)$  4 a. What is the value of

the coefficient x in this

equation? 40.07 g/mol b. What

is the molar mass of  $C_3H_4$ ? 2

mol O 2:1 mol H 2O

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Calculate the percentage yield  
2 60 mol of N<sub>2</sub> are mixed with  
120 mol of H<sub>2</sub> according ...

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**SHORT ANSWER** Answer the following questions in the



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space provided. 1. Given the following equation:  $C_3H_4(g) + x \cdot O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient  $x$  in this equation? b. What is the molar mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ? 2. a. What is meant by ideal conditions

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how many moles of  $\text{NH}_3$  will be  
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yield. 2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the ...

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Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry calculations start with a balanced chemical equation. This equation gives the ... 290 Chapter 9 DO NOT EDIT--Changes must be made through "File info" ...

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**Stoichiometry SECTION 3**

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**Stoichiometry SECTION 3**

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## ~~Stoichiometry Mixed Review Answers~~

### CHAPTER 9 REVIEW

### Stoichiometry CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER

Answer the following questions in the space provided.

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a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ?

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## **CHAPTER 9 REVIEW**

### **Stoichiometry MIXED REVIEW**

**SHORT ANSWER** Answer the

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mol O 2:1 mol H 20

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**PROBLEMS** Write the answer on the line to the left Show all your work in the space provided

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