

Chapter 18 Nuclear Chemistry Answer Key

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718 Chapter 18 Nuclear Chemistry + + + + Energy p + p + n + n ${}^{24}\text{He}^{2+}$ For many of the lighter elements, the possession of an equal number of protons and neutrons leads to stable atoms. For example, carbon-12 atoms, ${}^{12}\text{C}$ 12, with six protons and six neutrons, and oxygen-16 atoms, ${}^{16}\text{O}$ 16, with eight protons and eight neutrons, are both very stable.

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transmutation. the conversion of one element to another by nuclear reaction. nuclear fission. a nuclear reaction in which a nucleus splits into two or more lighter nuclei. chain reaction. a fission reaction in which the neutrons released initiate a second reaction, which in turn initiate a third reaction, and so on.

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Nuclear Decay Reactions. Just as we use the number and type of atoms present to balance a chemical equation, we can use the number and type of nucleons present to write a balanced nuclear equation for a nuclear decay reaction. This procedure also allows us to predict the identity of either the parent or the daughter nucleus if the identity of only one is known.

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protons and the number of neutrons, and Z, the atomic number, is the number of protons.

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transmutation. the conversion of one element to another by nuclear reaction. nuclear fission. a nuclear reaction in which a nucleus splits into two or more lighter nuclei. chain reaction. a fission reaction in which the neutrons released initiate a second reaction, which in turn initiate a third reaction, and so on.

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