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~~ACIDS AND~~

~~BASES: A BRIEF~~

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Answers

REVIEW • Acids and bases were first recognized by the properties of their aqueous solutions. o For example, acids turn litmus red, whereas bases turn litmus blue.

Chapter 16 Review
Acid Base Titration
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BASES. Table 16.1.

1: General

Properties of Acids
and Bases. produce
a piercing pain in a
wound. give a
slippery feel. taste
sour. taste bitter.

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Answers

are colorless when placed in phenolphthalein

(an indicator). are pink when placed in phenolphthalein (an indicator). are red on blue litmus paper (a pH indicator).

16.1: Acids and Bases - A Brief
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Answers
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Bronsted-Lowry
Acids and Bases;
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Autoionization of
Water; 16.4 The pH
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Acids and Bases;
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16.7 Weak Bases;
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Between K_a and K_b ; 16.9 Acid-Base
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What two substances are always produced by a neutralization reaction? a. acid and a base b. water and a base c. water and an acid d. water and a salt 2.

*Chapter 16:
Leader: Acid and
Base Review
Page 16/77*

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Acid Base
Titration Ph Mixed
Answers

Chapter 16 Acids and Bases 1. Acids were recognized primarily from their sour taste. Bases were recognized from their bitter taste and slippery feel on skin. 2. In the Arrhenius definition, an acid is a substance that produces hydrogen

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Acid Base
Titration Ph Mixed
Answers

ions (H^+) when dissolved in water, whereas a base is a substance that produces hydroxide ions (OH^-) in

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Chapter 16 Acids and Bases 16.1

Acids and Bases: A

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16.6 Weak Acids

16.7 Weak Bases

16.8 Relationship

between K_a and K_b

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The Arrhenius
definition is the

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narrowest view of
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...
Answers

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Base Equilibrium -
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Base ...*

16.1 Acids and
Bases: A Brief
Review • Acids taste
sour and cause
certain dyes to
change color.

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Acid Base
Titration Ph Mixed
Answers

- Bases taste bitter and feel soapy.
- Arrhenius concept

of acids and bases:

- An acid is a substance that, when dissolved in water, increases the concentration of H^+ ions.

- Example: HCl is an acid.
- An

Arrhenius base is a

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substance that,
when dissolved in
water, increases
the concentration
of OH^- ions.

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Base Equilibrium
Acids and Bases
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Titration, Ph, Mixed
Answers

acid-base reaction,
the equilibrium will
favor the reaction

that moves the
proton to the
stronger base. HCl
 $(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow$
 $\text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-$
 (aq) H_2O is a
much stronger base
than Cl^- , so the
equilibrium lies so
far to the right K is

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Titration Ph Mixed
Answers

not measured ($K \gg 1$).

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Behavior and

Chemical Structure

Inductive effects

and charge

delocalization

significantly

influence the

acidity or basicity

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Titration Ph Mixed
Answers

of a compound. The acid-base strength of a molecule

depends strongly on its structure.

The weaker the A-H or B-H⁺ bond, the more likely it is to dissociate to form an (H^+) ion.

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CHAPTER 16 - Acid-
Base Equilibria

Section 16.1 -

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Brief Review (a)

Define an acid and
a base, according
to the Arrhenius
definition. acid =
base = (b) Write

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Answers

the products of
each chemical
reaction below,

which involves the
dissociation of each
reactant into
aqueous ions.

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CHAPTER 16
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and bases were first recognized by the properties of their aqueous solutions. o For example, acids turn litmus red, whereas bases turn litmus blue.

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Answers
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Acids & Bases: A
Brief Review. -
Arrhenius acids
and bases: -- acid:
an H^+ donor $HA \rightleftharpoons H^+$
 $A(aq)$ (aq) (aq) --
base: an OH^- donor

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Acid Base
Titration Ph Mixed
Answers

$\text{MOH} + \text{M}^+ + \text{OH}^-(\text{aq})$
 $(\text{aq}) + (\text{aq}) -$

Brønsted-Lowry
acids and bases: --
acid: an H^+ donor
 $\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-(\text{aq})$
 (aq) -- base: an H^+
acceptor HB

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and bases were first recognized by the properties of their aqueous solutions. o For example, acids turn litmus red, whereas bases turn litmus blue.

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*K_a K_b K_w pH pOH
pK_a pK_b H⁺ OH⁻
Calculations - Acids*

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and bases were first recognized by the properties of their aqueous solutions. For example, acids turn litmus red, whereas bases turn litmus blue.

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and Bases. produce
a piercing pain in a
wound. give a
slippery feel. taste
sour. taste bitter.
are colorless when

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placed in phenolphthalein (an indicator). are pink when placed in phenolphthalein (an indicator). are red on blue litmus paper (a pH indicator).

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Acids and Bases;
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Water; 16.4 The pH
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16.6 Weak Acids;
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Between K_a and K_b ; 16.9 Acid-Base
Properties of Salt
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Structure; 16.11 ...

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Answers

substances are
always produced by
a neutralization
reaction? a. acid
and a base b. water
and a base c. water
and an acid d.
water and a salt 2.

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*Leader: Acid and
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and Bases 1. Acids were recognized primarily from their sour taste. Bases were recognized from their bitter taste and slippery feel on skin. 2. In the Arrhenius definition, an acid is a substance that produces hydrogen ions (H^+) when

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dissolved in water,
whereas a base is a
substance that
produces hydroxide
ions (OH^-) in

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16.7 Weak Bases
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The Arrhenius

definition is the

narrowest view of

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Answers
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Bases: A Brief
Review • Acids taste
sour and cause
certain dyes to
change color.

• Bases taste bitter

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and feel soapy.

• Arrhenius concept
of acids and bases:

• An acid is a substance that, when dissolved in water, increases the concentration of H^+ ions.

• Example: HCl is an acid. • An

Arrhenius base is a substance that,

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Answers

when dissolved in water, increases the concentration of OH^- ions.

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Acid and Base
Strength In any
acid-base reaction,
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Titration Ph Mixed
Answers

the equilibrium will favor the reaction that moves the

proton to the stronger base. $\text{HCl (aq)} + \text{H}_2\text{O(l)} \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ H_2O is a much stronger base than Cl^- , so the equilibrium lies so far to the right K is not measured (

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$K > 1$).

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Behavior and
Chemical Structure
Inductive effects
and charge
delocalization
significantly
influence the
acidity or basicity
of a compound. The

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Titration Ph Mixed
Answers

acid-base strength
of a molecule
depends strongly

on its structure.

The weaker the
A-H or B-H⁺ bond,
the more likely it is
to dissociate to
form an (H^+)
ion.

*16: Acid-Base
Equilibria -
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Base Equilibria

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Define an acid and
a base, according
to the Arrhenius
definition. acid =

base = (b) Write
the products of

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Answers

each chemical
reaction below,
which involves the
dissociation of each
reactant into
aqueous ions.

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Equilibria ...

Section 16.1 -

ACIDS AND

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BASES: A BRIEF REVIEW • Acids and bases were first recognized by the properties of their aqueous solutions. For example, acids turn litmus red, whereas bases turn litmus blue.

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Answers
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Acid-Base

Equilibria. 16.1

Acids & Bases: A

Brief Review. -

Arrhenius acids

and bases: -- acid:

an H^+ donor $HA \rightleftharpoons H^+$

$A(aq) \rightleftharpoons H^+(aq) + A^-(aq)$ --

base: an OH^- donor

$MOH \rightleftharpoons M^+ + OH^-(aq)$

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Titration Ph Mixed
Answers
(aq) (aq) -
Brønsted-Lowry
acids and bases: --

acid: an H⁺ donor

HA H A(aq) (aq)

(aq) -- base: an H⁺
acceptor HB

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chemical concepts
by starting with the
basics, using
symbols or
diagrams, and
conclude by
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BASES: A BRIEF REVIEW • Acids and bases were first recognized by the properties of their aqueous solutions. For example, acids turn litmus red, whereas bases turn litmus blue.

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Answers
*K_a K_b K_w pH pOH
pK_a pK_b H⁺ OH⁻
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Learners examine
an acid-base
titration graph and
answer four
questions about the
data.