

## Chapter 13 Chemical Kinetics

### Chapter 13 - Chemical Kinetics - Part I - Average and Instantaneous Rates

~~Chemical Kinetics Rate Laws – Chemistry Review – Order of Reaction \u0026 Equations~~  
~~Chemistry 102: Chapter 13 chemical kinetics (University of Jordan) || Part 1~~  
~~Chapter 14 – Chemical Kinetics: Part 1 of 17~~  
~~Chapter 5 (Gases) - Part 3 \u0026 Chapter 13 (Chemical Equilibrium) - Part 1~~  
~~Chapter 14 (Chemical Kinetics) – Part 1~~  
~~Chapter 14 Chemical Kinetics (L-13) Chemical Kinetics | 2nd order \u0026 nth order Rxn. | Half Life Calculation \u0026 1Q. By Arvind Arora~~  
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~~102~~  
~~12~~  
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~~Membrane structure and function | Part 1~~  
~~Reaction Rate Laws~~  
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### 13. Chemical Kinetics - Half Life Period

13.7 Kinetics of Catalysis • A catalyst is a substance that takes part in a chemical reaction and speeds it up but undergoes no permanent chemical change itself. • Catalysis can be classified into two types: homogeneous and heterogeneous.

### Chapter 13. Chemical Kinetics

Kinetics- Study of factors that affect how fast a reaction occurs and the step-by-step processes involved in chemical reactions. Factors that Affect Reaction Rate  
A. Concentration of reactants - higher reactant concentrations increase the rate of reaction.  
B. Catalyst – substance that accelerates the reaction rate without being transformed.

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participating in the slow or rate-determining step of a reaction mechanism.

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Chapter 13 Chemical Kinetics Student: \_\_\_\_\_ 1. The units of "reaction rate" are A.  $L \text{ mol}^{-1} \text{ s}^{-1}$ . B.  $L^2 \text{ mol}^{-2} \text{ s}^{-1}$ . C.  $\text{s}^{-1}$ . D.  $\text{s}^{-2}$ . E.  $\text{mol L}^{-1} \text{ s}^{-1}$ . 2. For the reaction  $\text{BrO}_3^- + 5\text{Br}^- + 6\text{H}^+ \rightarrow 3\text{Br}_2 + 3\text{H}_2\text{O}$  at a particular time,  $-\frac{d[\text{BrO}_3^-]}{dt} = 1.5 \times 10^{-2} \text{ M/s}$ .

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the factors affecting it and the mechanism by which the reactions proceed. 2. Rate of reaction is the change in concentration of reactants or products per unit time. For a general reaction,  $A+B \rightarrow C$ .

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chemical kinetics – area of chemistry dealing with speeds/rates of reactions. rates of reactions affected by four factors. concentrations of reactants. temperature at which reaction occurs. presence of a catalyst. surface area of solid or liquid reactants and/or catalysts.

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In Section 13.6 , you saw that it is possible to use kinetics studies of a chemical system, such as the effect of changes in reactant concentrations, to deduce events that occur on a microscopic scale, such as collisions between individual particles. Such studies have led to the collision model of chemical kinetics, which is a useful tool for understanding the behavior of reacting chemical species.

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